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Balancing by oxidation number method/Basic concepts and chemical calculation/ 11th Chemistry 2020

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Balance the following chemical equations: (a)  $\text{HNO}_3 + \text{Ca(OH)}_2?$

$\text{Ca(NO}_3)_2 + \text{H}_2\text{O}$  Balance the following chemical equations: 8th Class General Science - Ch 5 - Balancing the Chemical Equation- General Science 8th Class How to Balance any Equation in 2 minutes Balancing Chemical Equations in Hindi

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Equations L2 | CBSE Class 10 NCERT | Balancing of Chemical Equations | Umang Chemical Reactions \u0026 Equations | Balancing Chemical Reactions | Class 10th SSC Board Part 3 Chemfiesta Balance The Following Equations

Chemfiesta Balance The Following Equations Balancing Equations and Simple Stoichiometry-KEY Balancing Equations and Simple Stoichiometry-KEY Balance the following equations: 1)  $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$  2)  $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$  3)  $1 \text{ HBr} + 1 \text{ KHCO}_3$

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If you go through the process, you'll find that you end up with the following equation:  $\text{C}_2\text{H}_2\text{O} + 3 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O}$  For which the inventory leaves oxygen unbalanced

Balancing Equations | The Cavalcade o' Chemistry

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Equations Worksheet - Answers Note to students: It is acceptable to leave spaces blank when balancing equations - blank spaces are interpreted as containing the number "1".

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Reaction stoichiometry could be computed for a balanced equation. Enter either the number of moles or weight for one of the compounds to compute the rest. Limiting reagent can be computed for a balanced equation by entering the number of moles or weight for all reagents. Examples of complete chemical equations to balance:  $\text{Fe} + \text{Cl}_2 = \text{FeCl}_3$

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question 14 has coefficients of 2,3,3,1. Please email me if I am incorrect. berghmary@yahoo.com. Thanks again.

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<http://www.chemfiesta.com> Solutions for the Balancing Equations

Practice Worksheet 1)  $2 \text{NaNO}_3 + \text{PbO} \rightarrow \text{Pb}(\text{NO}_3)_2 + \text{Na}_2\text{O}$  2)  $6 \text{AgI} + \text{Fe}_2(\text{CO}_3)_3 \rightarrow 2 \text{FeI}_3 + 3 \text{Ag}_2\text{CO}_3$  3)  $\text{C}_2\text{H}_4\text{O}_2 + 2 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$  4)  $\text{ZnSO}_4 + \text{Li}_2\text{CO}_3 \rightarrow \text{ZnCO}_3 + \text{Li}_2\text{SO}_4$  5)  $\text{V}_2\text{O}_5 + 5 \text{CaS} \rightarrow 5 \text{CaO} + \text{V}_2\text{S}_5$  6)  $\text{Mn}(\text{NO}_2)_2 + \text{BeCl}_2 \rightarrow \text{Be}(\text{NO}_2)_2 + \text{MnCl}_2$  7)  $3 \text{AgBr} + \text{GaPO}_4 \rightarrow \text{Ag}_3\text{PO}_4 + \text{GaBr}_3$

## Balancing Equations Practice Worksheet

Balance the following equations and indicate the type of reaction taking place: 1)  $3 \text{NaBr} + 1 \text{H}_3\text{PO}_4 \rightarrow 1 \text{Na}_3\text{PO}_4 + 3 \text{HBr}$  Type of reaction: double displacement. 2)  $3 \text{Ca}(\text{OH})_2 + 1 \text{Al}_2(\text{SO}_4)_3 \rightarrow 3 \text{CaSO}_4 + 2 \text{Al}(\text{OH})_3$  Type of reaction: double displacement. 3)  $3 \text{Mg} + 1 \text{Fe}_2\text{O}_3 \rightarrow 2 \text{Fe} + 3 \text{MgO}$  Type of reaction: single displacement. 4)  $1 \text{C}_2\text{H}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$

## Balancing Equations Practice Worksheet

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(a) The O.N. (oxidation number) of P decreases from 0 in  $P_4$  to -3 in  $PH_3$  and increases from 0 in  $P_4$  to +2 in  $HPO_2$ . Hence,  $P_4$  acts both as an oxidizing agent and a reducing agent in this reaction.. Ion-



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electron method: The oxidation half equation is:  $\text{P}^{4+}(\text{s}) \rightarrow \text{H}_2\text{PO}_4^-(\text{aq})$   
. The P atom is balanced as:

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